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The video below refers to Dynamic Equilibrium (under the section Equilibrium Systems) 7.1 p. 420 in the Nelson Textbook The video below is related to 7.2 Equilibrium Law and the

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Equilibrium Constant (p.429
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Solution: OH OH OH 0.50 mmol
mL ?? ?=x = nc V ×25.0 mL OH
n ? =12.5 mmol HCl(aq)
HCl(aq) [HCl(aq)] 12.5 mmol
= = n V 0.10 mmol HCl(aq)
/mL V =120 mL 2. (a) Given:
V HCl(aq) = 25.00 mL;

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$[\text{HCl}(\text{aq})] = 0.350 \text{ mol/L}; V$

$\text{NaOH}(\text{aq}) = 5.0 \text{ mL};$

$[\text{NaOH}(\text{aq})] = 0.500 \text{ mol/L};$

Required: $[\text{H}^+(\text{aq})], \text{pH}$

Analysis: Before titration,

amount of H^+ is: $\text{H}(\text{aq}) \text{ H}(\text{aq})$

$\text{H}(\text{aq}) \quad 0.350 \text{ mmol} \quad \text{mL} \quad ++ \quad += \times =$

$\text{nc} \quad V$

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Section 8.7: Acid-Base
Titration Tutorial 1
Practice, page 547

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Solution: $Q = [H_2(g)] [I_2]$

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$$Q = \frac{[\text{HI}(\text{g})]^2}{[\text{H}_2(\text{g})][\text{I}_2(\text{g})]}$$

$$Q = \frac{(0.040)^2}{(0.010)(0.14)}$$

$Q = 0.020$ Statement: The value of Q is 0.020. Q is equal to K , so the reaction is at equilibrium. (b) Given:

$$[\text{HI}(\text{g})] = 0.20 \text{ mol/L}; [\text{H}_2(\text{g})] = 0.15 \text{ mol/L}; [\text{I}_2(\text{g})] = 0.15 \text{ mol/L}$$

$$Q = \frac{(0.20)^2}{(0.15)(0.15)} = 0.89$$

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= 0.090 mol/L; $K = 0.020$

Required: Q ; direction of
reaction to reach

equilibrium Solution: $Q = \frac{[\text{H}_2(\text{g})][\text{I}_2(\text{g})]}{[\text{HI}(\text{g})]^2} =$
 $\frac{(0.15)(0.090)}{(0.20)^2} Q = 0.34$

Section 7.5: Quantitative

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Changes in Equilibrium
Systems ...

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